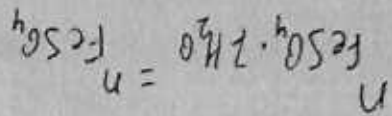


Umkehr: 868,56 T

$$m_{\text{FeSO}_4 \cdot 7\text{H}_2\text{O}} = 3,29 \cdot (152 + 7 \cdot 16) = 868,56 \text{ T}$$

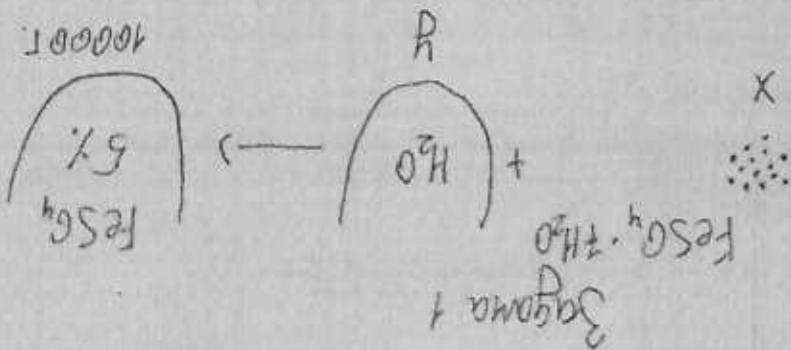


$$n_{\text{FeSO}_4} = \frac{500}{152} = 3,29 \text{ Mol}$$

8 packung

$$m_{\text{FeSO}_4} = 10000 \cdot 0,05 = 500 \text{ T}$$

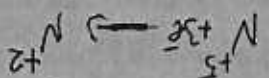
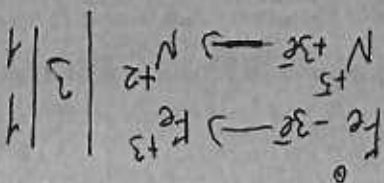
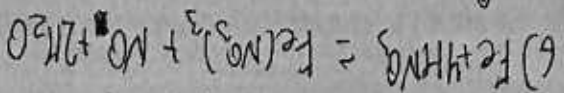
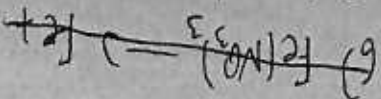
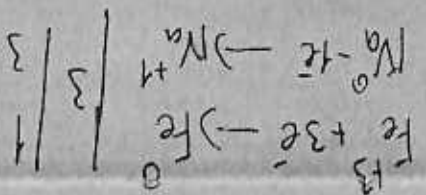
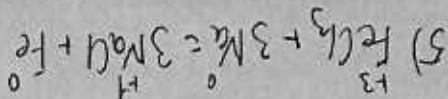
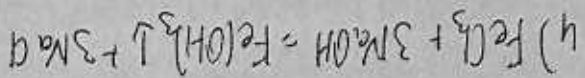
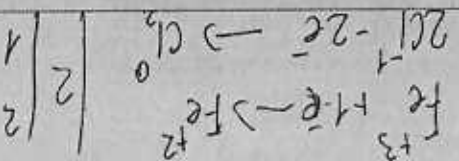
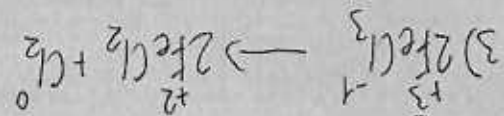
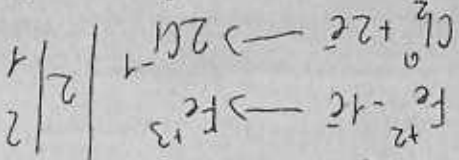
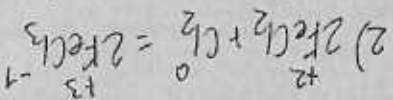
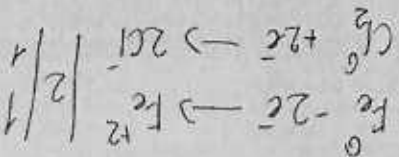
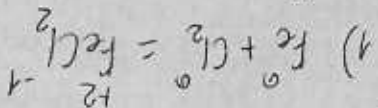
$$m_{\text{p-pa FeSO}_4} = 10000 \cdot 1 = 10000 \text{ T}$$

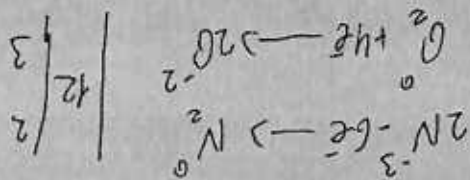
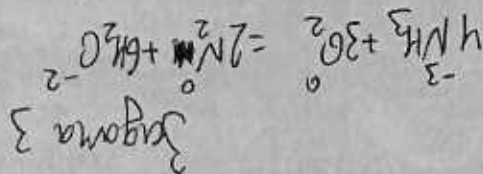


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3. Aufgabe

1) $X - Fe$





$$V_{\text{NH}_3}^{\text{норм}} = 100 \cdot 0,75 = 75 \text{ л}$$

$$n_{\text{NH}_3} = \frac{75}{22,4} = 3,348 \text{ моль}$$

$$V_{\text{O}_2}^{\text{норм}} = 80 \cdot 0,9 = 72 \text{ л}$$

$$n_{\text{O}_2} = \frac{72}{22,4} = 3,214 \text{ моль}$$

$$n_{\text{N}_2} = \frac{1}{2} \cdot n_{\text{NH}_3}$$

$$n_{\text{N}_2} = \frac{3,348}{2} = 1,674 \text{ моль}$$

$$V_{\text{N}_2} = 1,674 \cdot 22,4 = 37,49 \text{ л}$$

$$V_{\text{N}_2}^{\text{норм}} = 37,49 \cdot 0,95 = 35,62 \text{ л}$$

200

Задача 4

$$f_a - 930 \text{ r.}$$

$$\text{тога} - x \text{ r.}$$

$$x = 9300 \text{ r.}$$

$$m_p = 9300 \text{ r.}$$

$$n_p = \frac{9300}{31} = 300 \text{ молекул}$$

$$n_{NH_4H_2PO_4} = n_p$$

$$n_{NH_4H_2PO_4} = 300 \text{ молекул}$$

$$m_{NH_4H_2PO_4} = 300 \cdot 115 = 34500 \text{ r.} = 34,5 \text{ кг.}$$

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Задача 5

Условие:

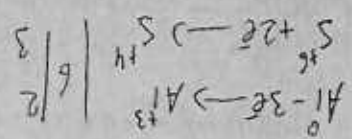
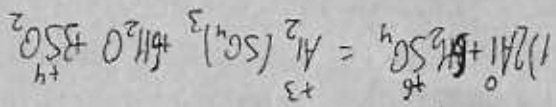
Al

ZnO

KOH

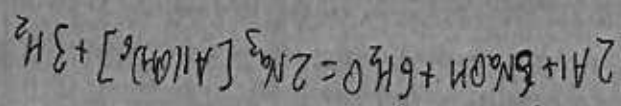
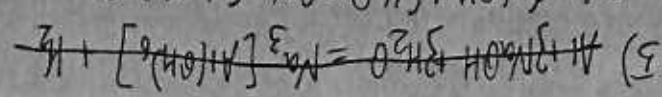
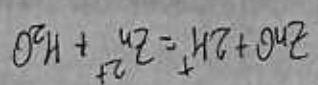
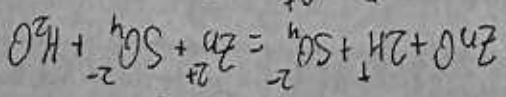
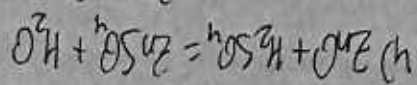
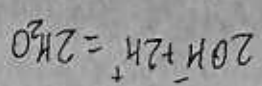
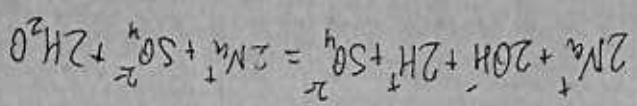
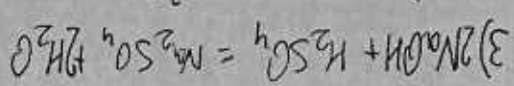
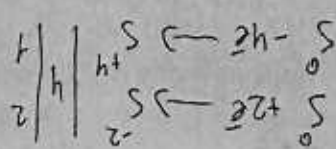
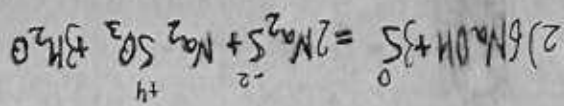
S

H₂SO₄



Al⁰ - восстановитель

H₂SO₄(S⁺⁶) - окислитель



100

Задача 5

Уравнение

