

$$m_p \cdot p_a = p_r \cdot p_a \cdot V_r \cdot p_a = 1 \cdot 10 \cdot 1000 = 10000 \text{ z} = 10 \text{ kN}$$

$$m(\text{FeSO}_4) = 0,05 \cdot m_p \cdot p_a = 500 \text{ z}$$

$$n(\text{FeSO}_4) = \frac{m(\text{FeSO}_4)}{M(\text{FeSO}_4)} = \frac{500}{56+32+16 \cdot 4} \approx 3,28 \text{ mola}$$

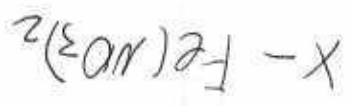
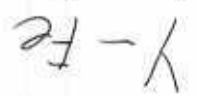
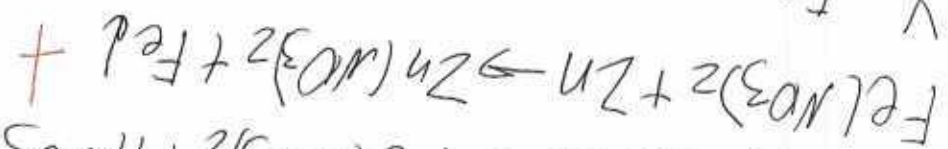
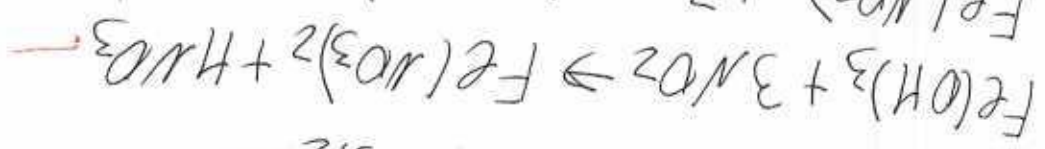
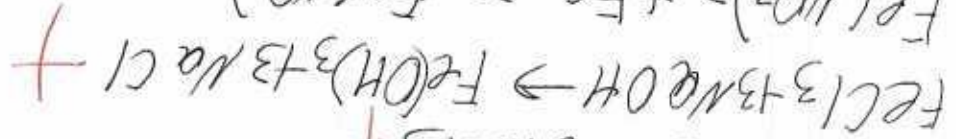
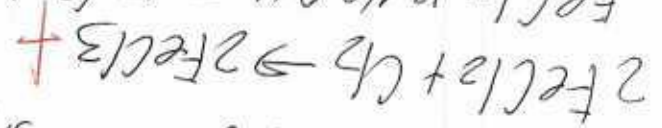
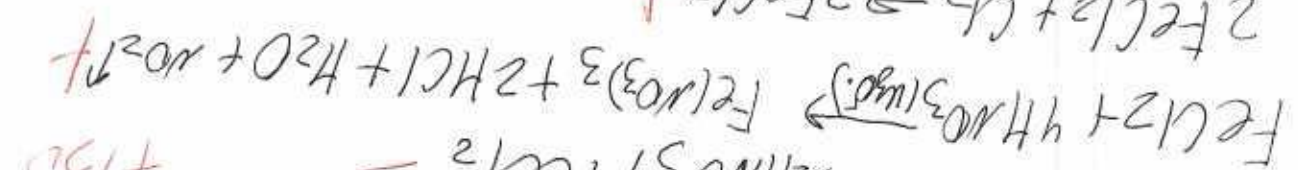
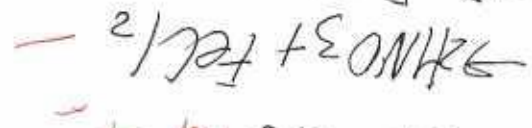
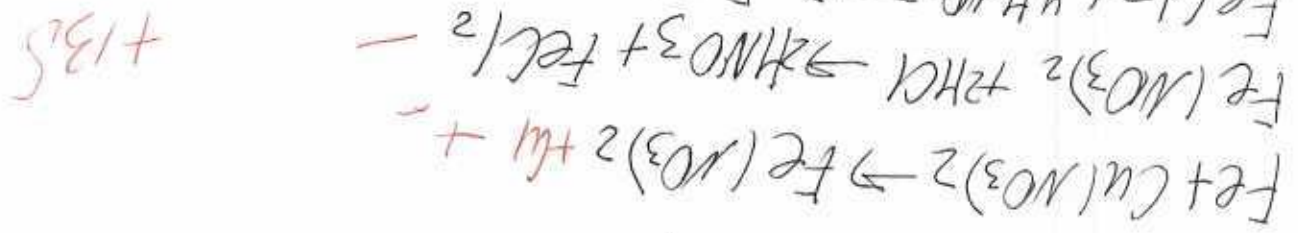
$$n(\text{FeSO}_4) = n(\text{FeSO}_4 \cdot \frac{1}{4} \text{H}_2\text{O}) = 3,28 \text{ mola}$$

$$m(\text{FeSO}_4 \cdot \frac{1}{4} \text{H}_2\text{O}) = n(\text{FeSO}_4 \cdot \frac{1}{4} \text{H}_2\text{O}) \cdot M(\text{FeSO}_4 \cdot \frac{1}{4} \text{H}_2\text{O})$$

$$m(\text{FeSO}_4 \cdot \frac{1}{4} \text{H}_2\text{O}) = 3,28 \cdot (56+32+16 \cdot 4 + \frac{1}{4} \cdot (2+16)) = 974,5 \text{ z}$$

$$\text{Omben: } m(\text{FeSO}_4 \cdot \frac{1}{4} \text{H}_2\text{O}) = 974,5 \text{ z} + 205$$

N/2



1/4



$$V_{\text{H}_2\text{O}} = 80 \text{ l} \cdot (1 - 0,1) = 72 \text{ l} \rightarrow \text{H}_2\text{O} \text{ - H}_2\text{O} \text{ - H}_2\text{O}$$

$$10\% = 0,1$$

III. k. Van y zayob obagayon naypabawon, mo lbtet mawet  
 ut coemotum

H<sub>2</sub> nayayaw:

$$\frac{n(\text{O}_2)}{n(\text{H}_3)} = \frac{4}{3} = \frac{V_1(\text{O}_2)}{V_1(\text{H}_3)} = 0,42$$

$$\frac{V_{\text{H}_2\text{O}}}{V_{\text{H}_2}} = \frac{42}{70} = 0,42$$

III. k. O<sub>2</sub> b nayayawet.

IV. k.:

~~IV. k.:~~

$$n(\text{H}_2) = \frac{3}{2} \cdot V_{\text{H}_2\text{O}} = 48 \text{ l}$$

H<sub>2</sub> - za dawaga 95%:

$$n_1(\text{H}_2) = n(\text{H}_2) \cdot 0,95 = 45,6 \text{ l}$$

ut H<sub>2</sub> - za nayayawet:

$$n_2(\text{H}_2) = n_1(\text{H}_2) \cdot (1 - 0,25) = 34,2 \text{ l}$$

$$25\% = 0,25$$

+

+165

$$\text{Emblem: Odygyn } n_2 = 34,2 \text{ l}$$

$$n(P) = \frac{m(P)}{M(P)} = \frac{930}{34} = 30 \text{ mol}$$

$$n(NH_4H_2PO_4) = n(P) = 30 \text{ mol}$$

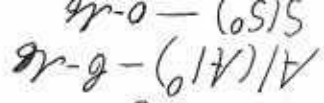
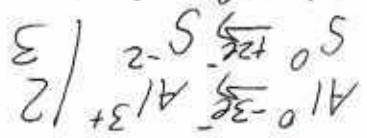
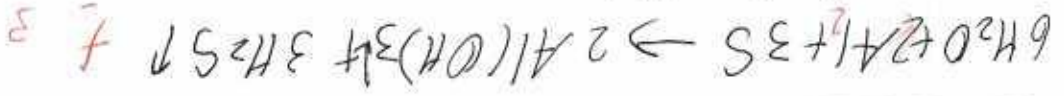
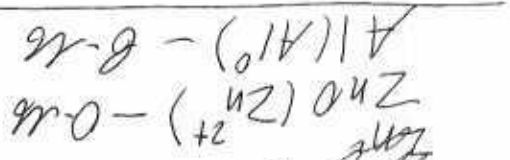
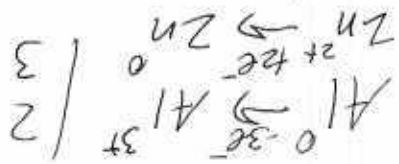
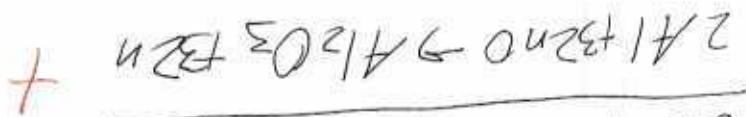
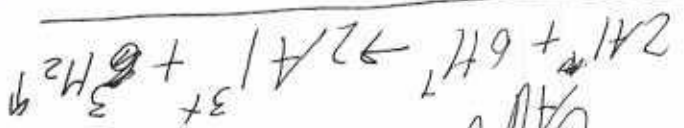
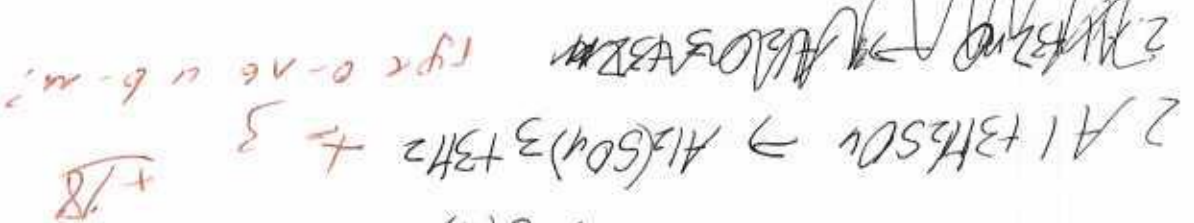
$$m(NH_4H_2PO_4) = n(NH_4H_2PO_4) \cdot M(NH_4H_2PO_4) = 30 \cdot (14 + 4 + 31 + 76.9) = 30 \cdot 125.9 = 3777 \text{ g}$$

$$= 3777 \text{ g} - 345 \text{ g} = 3432 \text{ g}$$

$$m_2(NH_4H_2PO_4) = 10 \cdot m(NH_4H_2PO_4) = 10 \cdot 345 = 3450 \text{ g}$$

$$m_2(NH_4H_2PO_4) = 3450 \text{ g} + 200 \text{ g} = 3650 \text{ g}$$

$$N_5(1)$$



$$3/4$$

h/h

